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UTILITY PATENT APPLICATION TRANSMITTAL Only for new nonprovisional applications under 37 CFR 1.53(b)	Attorney Docket No.:	1856-00301
	First Inventor or Application Identifier	Alfred E. Keller
	Title	Process for Producing Syngas in a Short Contact Time Reactor Using Catalytic Partial Oxidation of Hydrogen Sulfide
	Express Mail Label No.	EL497839799US

APPLICATION ELEMENTS See MPEP chapter 600 concerning utility patent application contents.	ADDRESS TO: Assistant Commissioner for Patents Box Patent Application Washington, DC 20231
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<p>1. <input checked="" type="checkbox"/> Fee Transmittal Form (e.g. PTO/SB17) (Submit an original, and a duplicate for fee processing)</p> <p>2. <input checked="" type="checkbox"/> Specification Total Pages 16 (preferred arrangement set forth below) - Descriptive title of the invention - Cross References to Related Applications - Statement Regarding Fed sponsored R & D - Reference to Microfiche Appendix - Background of the Invention - Brief Summary of the Invention - Brief Description of the Drawings (if filed) - Detailed Description - Claims(s) - Abstract of the Disclosure</p> <p>3. <input checked="" type="checkbox"/> Drawing(s) (35 USC 113) Total Sheets 2</p> <p>4. <input checked="" type="checkbox"/> Oath or Declaration Total Pages 1 a. <input checked="" type="checkbox"/> Newly executed (original or copy) b. <input type="checkbox"/> Copy from a prior application (37 CFR 1.63(d)) (for continuation/divisional with Box 17 completed) i. <input type="checkbox"/> <u>DELETION OF INVENTOR(S)</u> Signed statement attached deleting inventor(s) named in the prior application, see 37 CFR 1.63(d)(2) and 1.33(b).</p>	<p>5. <input type="checkbox"/> Microfiche Computer Program (Appendix)</p> <p>6. <input type="checkbox"/> Nucleotide and/or Amino Acid Sequence Submission (if applicable, all necessary) a. <input type="checkbox"/> Computer Readable Copy b. <input type="checkbox"/> Paper Copy (identical to computer copy) c. <input type="checkbox"/> Statement verifying identity of above copies</p> <p>ACCOMPANYING APPLICATION PARTS</p> <p>7. <input checked="" type="checkbox"/> Assignment Papers (cover sheet & document(s))</p> <p>8. <input type="checkbox"/> 37 CFR 3.73(b) Statement [] Power of Attorney (when there is an assignee)</p> <p>9. <input type="checkbox"/> English Translation Document (if applicable)</p> <p>10. <input type="checkbox"/> Information Disclosure Statement (IDS)/PTO-1449 <input type="checkbox"/> Copies of IDS Citations</p> <p>11. <input type="checkbox"/> Preliminary Amendment</p> <p>12. <input checked="" type="checkbox"/> Return Receipt Postcard (MPEP 503) (Should be specifically itemized)</p> <p>13. <input type="checkbox"/> Small Entity Statement(s) <input type="checkbox"/> Statement filed in prior application. Status still proper and desired</p> <p>14. <input type="checkbox"/> Certified Copy of Priority Document(s) (if foreign priority is claimed)</p> <p>15. <input checked="" type="checkbox"/> Other Power of Attorney by Assignee</p>
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16. If a CONTINUING APPLICATION, check appropriate box and supply the requisite information.
☐ Continuation ☐ Divisional ☐ Continuation-in-part (CIP) of prior application No.: ____/_____
Prior application information: Examiner: * Group/Art Unit: *
For CONTINUATION or DIVISIONAL APPS only: The entire disclosure of the prior application, from which an oath or declaration is supplied under box 4b, is considered a part of the disclosure of the accompanying continuation or divisional application and is hereby incorporated by reference. The incorporation can only be relief upon when a portion has been inadvertently omitted from the submitted application parts.

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**PROCESS FOR PRODUCING SYNGAS IN A SHORT CONTACT TIME REACTOR
USING CATALYTIC PARTIAL OXIDATION OF HYDROGEN SULFIDE**

CROSS-REFERENCE TO RELATED APPLICATIONS

This application claims the benefit of provisional application no. 60/146,635 filed July 30, 1999, the disclosure of which is incorporated herein by reference.

STATEMENT REGARDING FEDERALLY SPONSORED
RESEARCH OR DEVELOPMENT

Not applicable.

BACKGROUND OF THE INVENTION

Technical Field of the Invention

The present invention generally relates to processes for the catalytic oxidation of light hydrocarbons to synthesis gas (syngas), and more particularly to methods of increasing the yield of syngas in processes employing partial oxidation of methane or natural gas to products containing CO and H₂. Still more particularly, the processes of the present invention relate to such methods in which the concurrent oxidation reaction (i.e., the complete combustion of methane to carbon dioxide and water) is replaced with combustion of hydrogen sulfide to sulfur and water, to improve the yield of synthesis gas.

Description of Related Art

Many refineries face an abundant supply of lower alkanes, i.e., C₁-C₄ alkanes such as methane, and relatively few means of converting them to more valuable products. Much research has been devoted to investigating the conversion of methane to more easily transportable products. One technique that has been well-developed entails the partial oxidation

of light hydrocarbons in the presence of a catalyst. This technique results in the production of synthesis gas, *i.e.*, “syngas”, a mixture of carbon monoxide and hydrogen. The catalytic partial oxidation of methane can be represented by the following reaction scheme:



5 At the same time, some of the methane burns completely, according to the equation:



Hence, syngas is typically a mixture of carbon monoxide and molecular hydrogen, generally having a hydrogen to carbon monoxide molar ratio in the range of 1:5 to 5:1, and which may contain other gases such as carbon dioxide.

10 Synthesis gas has utility as a feedstock for conversion to alcohols, olefins, or saturated hydrocarbons (paraffins) according to the well-known Fischer-Tropsch process, and by other means. Synthesis gas is not a commodity; instead, it is typically generated on-site for further processing. The uses for syngas include, but are not limited to, a feedstock for conversion to high molecular weight (*e.g.* C₅₀₊) paraffins, which in turn provide an ideal feedstock for
15 hydrocracking, a feedstock for conversion to high quality jet fuel, and superior high octane value diesel fuel blending components. Another potential application of synthesis gas is for large-scale conversion to methanol.

Syngas generation from methane typically takes place by one of three techniques. Steam reforming of methane is the most common, followed by partial oxidation, and autothermal
20 reforming. Emerging technologies that have been developed to generate syngas from methane include a technique that entails exposing a mixture of methane and oxygen to a hot catalyst for a brief time, typically on the order of milliseconds or less, followed by cooling of the resultant gas stream. EP 303,438 describes a process for synthesis gas production by catalytic partial oxidation

to overcome some of the disadvantages and costs of steam reforming. A monolith catalyst is used with or without metal addition to the surface of the monolith and the process operates at space velocities of 20,000-500,000 hr⁻¹. Conventional catalytic partial oxidation processes are also described in U.S. Patent Nos. 5,654,491 and 5,639,929, the disclosures of which are
5 incorporated herein by reference.

Although in conventional syngas generation systems the syngas reaction is self-sustaining once initiated, it has been shown that 10 - 15 % of the carbon initially present as methane can be lost to the formation of CO₂ in combustion via equation (2) above. This directly reduces the yield of syngas that can be obtained. Hence, it is desirable to provide a syngas generation system
10 that allows a better yield of carbon monoxide and hydrogen.

In a related aspect of petroleum refining, many petroleum feed streams and separated fractions contain sulfur. Sulfur is typically undesirable in most petroleum refining processes and products. Therefore, refineries typically upgrade the quality of the various petroleum fractions by removing the sulfur. Specifically, hydrodesulfurization units are used to break down the
15 sulfur compounds in the petroleum fractions and convert the sulfur to H₂S. In addition to hydrodesulfurization processes, other conversion processes in a typical refinery, such as fluid catalytic cracking, coking, visbreaking, and thermal cracking, produce H₂S from sulfur containing petroleum fractions. The H₂S from both the desulfurization processes and these
20 conversion processes is typically removed from the gas streams or light liquid hydrocarbon streams using either chemical solvents based on alkanolamine chemistry or physical solvents. A circulating, regenerative H₂S removal system employing an absorption stage for H₂S pickup and a regeneration stage for H₂S rejection produces a concentrated stream of H₂S.

gas stream with a hot catalyst to form a product stream, and removing syngas and elemental sulfur from the product stream.

The present invention also provides a method for improving the yield of a syngas generation system, comprising providing a first gas stream comprising a light hydrocarbon, mixing a second gas stream comprising H_2S with the first gas stream to form a feed gas stream, while maintaining said feed gas stream below 500 degrees C, contacting the feed gas stream with a hot catalyst to form a product stream wherein less than 10% of the light hydrocarbon is converted to carbon dioxide, and removing syngas and elemental sulfur from the product stream.

BRIEF DESCRIPTION OF THE DRAWINGS

For a more detailed description of the present invention, reference will now be made to the accompanying drawings, wherein:

Fig. 1 is an enlarged cross-section of a reactor constructed in accordance with a preferred embodiment; and

Fig. 2 is a schematic diagram of the components of one preferred embodiment of the present system including the reactor of Fig. 1.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

Several schemes for carrying out partial oxidation are known in the art. As discussed above, one scheme for carrying out the exothermic oxidation reaction entails a brief exposure of the methane feed to a hot catalyst followed by cooling the resultant gas stream. The hot catalyst is positioned in the flow path of the feed gas. The catalyst comprises a wire gauze, several layers of wire gauze, or a porous ceramic. The catalyst is designed so that only a first fraction of the

feed gas contacts the catalyst, while the balance of the gas serves to quickly cool the first fraction and prevent the oxidation reaction from proceeding too far.

According to the present invention, a stream of H₂S is added to the feed stream of the syngas generation system. The feed stream comprises methane or a similar light hydrocarbon.

- 5 Pure oxygen is preferably mixed with the gas feed immediately before contacting the feed gases with a catalyst. Air, or a mixture of air and oxygen can be substituted for the pure oxygen. The H₂S reacts with oxygen in the stream according to the reaction:



- According to the present invention, H₂S partial oxidation is incorporated into the syngas generation scheme as follows. Referring initially to Figure 1, a preferred embodiment of the present system includes a reactor 10 that includes feed injection openings 12, 14, and 16, a mixing zone 19, a reaction zone 20 and a cooling zone 30. Reaction zone 20 preferably includes a thermal radiation barrier 22 positioned immediately upstream of a catalytic device 24. Radiation barrier 22 is preferably a porous ceramic or refractory material that is suited to
10 withstand operating temperatures and provide sufficient thermal insulation, such as are described in U.S. Patent 4,038,036 to Beavon, which is incorporated herein by reference in its entirety.

- Catalytic device 24 is preferably a layer or layers of wire gauze 25 or a porous ceramic monolith (not shown) having a suitable catalyst supported on its surface. Gauze 25 is preferably one or more layers of a substantially planar, flexible woven metal-containing or metal-coated
20 screen or gauze having about 20-120 mesh. More preferably, it is a gauze of metal wires about 25 micrometers to about 2.5 millimeters in diameter, which are made of about 87-93% by weight (wt-%) Pt and about 7-13 wt-% Rh. Alternative catalyst structures could include a disk with multiple perforations formed there through, a honeycomb-like structure, an etched foil and any

other structure that provides the desired amount of transparency to effect the desired partial oxidation. A detailed discussion of the catalyst structure and composition can be found in U.S. Patent No. 5,654,491 to Goetsch *et al.*, which is incorporated herein in its entirety.

Examples of suitable catalysts that can be included in the metal of the gauze or incorporated at its surface include, but are not limited to, platinum, rhodium, iridium, nickel, palladium, iron, cobalt, rhenium, rubidium, Pd-La₂O₃, Pt/ZrO₂, Pt/Al₂O₃.

In operation, a light hydrocarbon, such as methane, is fed into one of the feed injection openings 12. H₂S is fed into a second feed injection opening 14. Air or oxygen is fed into the third feed injection opening 16, which is preferably positioned close to catalyst 24. It will be understood that the feed injection openings can be configured differently from the configuration shown without affecting the principles or operation of the present system.

As the feed gases from feed injection openings 12, 14, 16 flow toward catalytic device 24, they are preferably subjected to thorough mixing by static mixer 18. During mixing, they are shielded by radiation barrier 22 from radiant heat that is generated downstream in the process. It is preferred that the temperature on the upstream side of barrier 22 be in the range of about 20°C to about 300°C. The feed gas stream is preferably at ambient temperature prior to contact with the catalyst. Preheating the feed gas stream is not desired, as it can cause homogeneous reactions and reduce the selectivity of the process of the present invention for the desired compounds. Therefore, preheating the feed gas mixture is typically avoided, although in some applications feed gas temperatures up to about 300° C can be tolerated.

After the gases pass barrier 22, they flow past catalytic device 24 and are simultaneously heated to an oxidation temperature in the range of from about 900 °C to about 1300 °C. The gas flow rate is preferably maintained such that the contact time for the portion of the gas that

contacts the catalyst is from about .001 to .01 seconds and more preferably from about .001 to .005 seconds.

This degree of contact produces a favorable balance between competing reactions and produces sufficient heat to maintain the catalyst at the desired temperature. Specifically, sulfur is produced by catalyzed partial oxidation according to equation (3) above, where x equals 2, 6, or 8, with x = 2 being the most likely. At the same time, exposure to the hot catalyst and oxygen partially oxidizes the hydrocarbons in the feed, according to the equation:



Oxygen for these reactions comes from the air, oxygen, or air/oxygen mix that is fed into the system with the H₂S and hydrocarbon feed gases.

Typically, the catalyst structure is heated as a result of the exothermic chemical reactions occurring at its surface; however, it can additionally or alternatively be heated by external means, such as electrical resistance, magnetic induction, RF, etc. Heating by external means can allow for increases in the rate at which feed gas can be passed through the catalyst structure while still obtaining desirable reaction products. In many cases it is helpful to heat the catalytic device with external means at least at the start of the process, so as to initiate the exothermic reactions on the catalyst structure. This initial heating can be accomplished in any suitable manner including electrical resistance, magnetic induction, RF, or the like. Once the system is running, it is preferably is run adiabatically or nearly adiabatically (i.e., without loss of heat), so as to reduce the formation of carbon (e.g., coke) on the surface of the gauze catalyst.

According to the present invention, the rate of feed of H₂S into the system is controlled and adjusted so that the heat generated by the oxidation of the H₂S is sufficient to maintain the desired temperature in reaction zone 20 and thus reduce the amount of the light hydrocarbon that

is completely combusted. Hence, the mole ratio of H₂S to light hydrocarbon in the feed is preferably in the range of from about 1:10 to about 2:3. Where the light hydrocarbon is methane, a preferred ratio of H₂S to methane is 2:3.

The rapid heating of the feed gases as a result of contact with the hot catalyst promotes fast reaction rates. In accordance with the present invention, the feed gas stream velocity past catalyst structure 10 is preferably at least about 0.1 meter/second, often as high as 4-5 meters/second, and even as high as 70 meters/second. The maximum velocity will generally be determined by the specific equipment used; however, the theoretical limit is that velocity at which the reaction would be extinguished. If an external means of heating the catalytic device 24 is used, this theoretical limit is significantly large.

According to one preferred embodiment, the feed gas stream velocity is between about 0.1 and 100 meters/second. As a result, the superficial contact time of the feed gas stream with a preferred embodiment of gauze catalytic device 24 is less than about 10,000 microseconds, and typically within a range of about 1000-10,000 microseconds. When used in the present invention, it is preferred that the superficial contact time of the feed gas stream with the catalyst be less than about 5000 microseconds, more preferably less than about 2000 microseconds. Superficial contact time is inversely proportional to the term "space velocity" that is used in many chemical process descriptions.

Although for ease in comparison with prior art, space velocities at standard conditions have been used to describe the present invention, it is well recognized in the art that residence time is the inverse of space velocity and that the disclosure of high space velocities equates to low residence times.

From reaction zone 20, the reacted gases preferably enter a firetube boiler 40, where they are cooled to below 450 °C and preferably to below 340 °C. As shown, it is preferred that heat removed from the partially oxidized gases can be recaptured in steam heating or the like. The rapid cooling that occurs in the boiler drops the temperature to below about 450° C and thus ceases the above reactions. A detailed description of the considerations involved in operating a reactor using millisecond contact times is given in U.S. Patent No. 5,654,491, which is incorporated herein in its entirety.

Accordingly, and referring now to Figure 2, the present system preferably includes the synthesis gas reactor 10, firetube boiler 40, a condenser 50, heater 55, a tailgas cleanup units 60, a cooler, and a quench tower 80. The cooled, partially oxidized gases flow from boiler 40 into condenser 50, where they are cooled further until the dew point of the elemental sulfur is reached. This allows for the removal of elemental sulfur, as desired, from the process. Once the bulk of the elemental sulfur is removed, the partially oxidized gases are reheated in heater 55 and passed through a tailgas converter unit 60. Each tailgas converter unit 60 includes at least a sulfur absorbing material 56 in contact with the fluid. More specifically, in each converter unit 60, the hot gas stream is passed over a bed of zinc or iron oxide. In this bed, any elemental sulfur is converted to metal sulfide and retained in the bed.

The effluent from the sulfur absorber is then preferably cooled sufficiently to condense the bulk of any remaining water from the gas stream.

The treated gases, which comprise CO, hydrogen, and nitrogen, are then sent to a synthesis gas conversion unit such as a methanol plant or a Fischer-Tropsch plant.

By substituting the heat of the partial oxidation of H₂S for the combustion of methane, the present invention provides the heat necessary to maintain the syngas reaction at the desired

temperature without giving up the methane to combustion products. This, in turn, results in a higher overall yield for the process.

While a preferred embodiment of the present invention has been shown and described, it will be understood that variations can be to the preferred embodiment, without departing from the scope of the present invention. For example, the mixing process can be altered or replaced with an active mixer, the thermal barrier can be modified, the structure and composition of the catalyst can be varied, and the tail gas treatment steps can be modified.

The complete disclosure of all patents, patent documents, and publications cited herein are incorporated by reference. The foregoing detailed description and examples have been given for clarity of understanding only. No unnecessary limitations are to be understood therefrom. The invention is not limited to the exact details shown and described, for variations obvious to one skilled in the art will be included within the invention defined by the claims.

CLAIMS

What is claimed is:

1. A method for improving the yield of a syngas generation system, comprising providing a first gas stream containing a light hydrocarbon, mixing a second gas stream containing
5 H₂S with the first gas stream to form a feed gas stream, mixing the feed gas stream with an oxygen containing stream to form a mixed feed stream, contacting the mixed feed stream with a hot catalyst to form a product stream, and removing syngas and elemental sulfur from the product stream.
- 10 2. The method according to claim 1 further comprising removing residual H₂S from the product stream.
3. The method according to claim 1 wherein mixing a second gas stream comprising H₂S with the first gas stream to form a feed gas stream is carried out at temperatures below
15 500 degrees C.
4. The method according to claim 1 wherein contacting the feed gas stream with a hot catalyst to form a product stream is carried out at temperatures above 500 degrees C.
- 20 5. The method according to claim 1 wherein less than 10% of the light hydrocarbon is converted to carbon dioxide.

6. The method according to claim 1 wherein the catalyst contact time is less than .01 seconds.

7. The method according to claim 1 wherein the catalyst is selected from the group consisting of: platinum, rhodium, iridium, nickel, palladium, iron, cobalt rhenium rubidium, Pd-La₂O₃, Pt/ZrO₂, Pt/Al₂O₃ and combinations thereof.

8. A system for the partial oxidation of light hydrocarbons, comprising a hydrocarbon injection line, an H₂S injection line in communication with said hydrocarbon injection line, an oxygen injection line in communication with said hydrocarbon injection line, a reaction zone receiving gases from said hydrocarbon, H₂S and oxygen injection lines and including a catalyst suitable for catalyzing said hydrocarbon to form CO and H₂.

9. The system according to claim 8 comprising a mixing zone upstream of said reaction zone, said mixing zone receiving gases from said hydrocarbon and said H₂S lines, wherein the temperature of said mixing zone is less than 500 degrees C.

10. The system according to claim 9 comprising a thermal barrier between said mixing zone and said reaction zone.

11. The system according to claim 9 wherein said oxygen injection line communicates with said reaction zone.

12. The system according to claim 9 wherein said mixing zone receives oxygen from said oxygen injection line.

13. The system according to claim 8 wherein the temperature of said reaction zone is greater than 500 degrees C.

14. The system according to claim 8 comprising at least one cooling zone downstream of said reaction zone.

15. The system according to claim 14 comprising at least one tailgas processing unit downstream of said cooling zone.

16. The system according to claim 8 wherein said catalyst is supported on a wire gauze.

17. The method according to claim 8 wherein the catalyst is selected from the group consisting of: platinum, rhodium, iridium, nickel, palladium, iron, cobalt rhenium rubidium, Pd-La₂O₃, Pt/ZrO₂, Pt/Al₂O₃ and combinations thereof.

18. A method for improving the yield of a syngas generation system, comprising providing a first gas stream comprising a light hydrocarbon, mixing a second gas stream comprising H₂S with the first gas stream to form a feed gas stream, while maintaining said feed gas stream below 500 degrees C, contacting the feed gas stream with a hot catalyst to form a

product stream wherein less than 10% of the light hydrocarbon is converted to carbon dioxide, and removing syngas and elemental sulfur from the product stream.

19. The method according to claim 18 comprising mixing oxygen with the light hydrocarbon
5 prior to contacting the feed gas stream with a hot catalyst.

20. The method according to claim 18 comprising mixing oxygen with the light hydrocarbon
during the contacting of the feed gas stream with a hot catalyst.

ABSTRACT OF THE DISCLOSURE

A method and system for improving the yield of syngas from the partial oxidation of methane or other light hydrocarbons is disclosed. The increase in yield results from the substitution of H_2S partial oxidation for the combustion of light hydrocarbon to CO_2 and water.

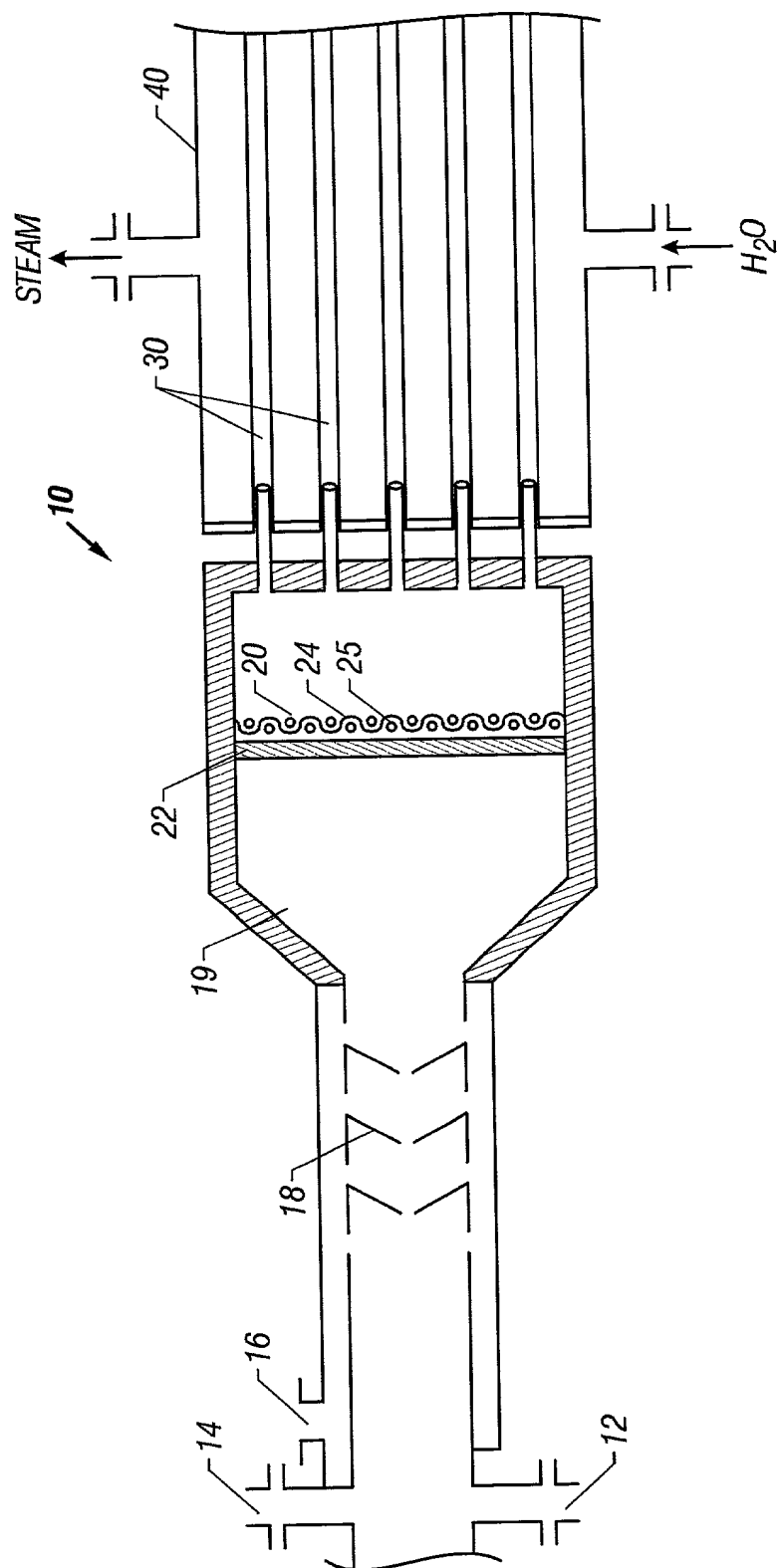


FIG. 1

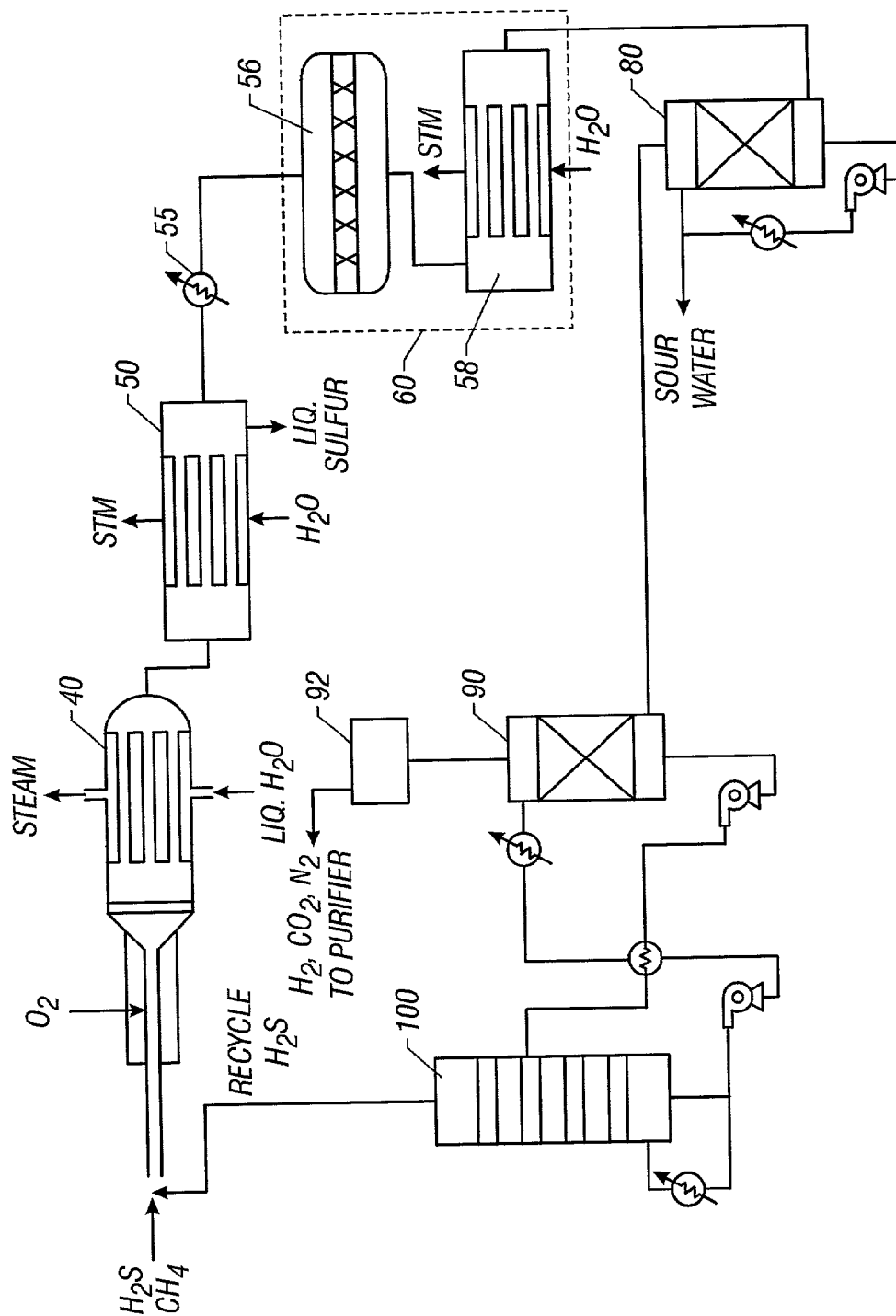


FIG. 2

DECLARATION**SOLE/JOINT INVENTOR
ORIGINAL/SUBSTITUTE/CIP**

As a below named inventor, I hereby declare that: my residence, post office address, and citizenship are as stated below next to my name. I believe I am the original, first, and sole inventor (if only one name is listed below) or a joint inventor (if plural inventors are listed below) of the subject matter which is claimed and for which a patent is sought on the invention entitled: *Process For Producing Syngas in a Short Contact Time Reactor Using Catalytic Partial Oxidation of Hydrogen Sulfide* described in the attached patent Application Serial No. _____, filed _____.

I hereby state that I have reviewed and understand the contents of the above identified specification, including the claims, as amended by any amendment referred to above; that I do not know and do not believe the same was ever known or used in the United States of America before my or our invention thereof, or patented or described in any printed publication in any country before my or our invention thereof or more than one year prior to this application; that the invention has not been patented or made the subject of an inventor's certificate issued before the date of this application in any country foreign to the United States of America on an application filed by me or my legal representative or assigns more than twelve months prior to this application; and that I acknowledge the duty to disclose information of which I am aware which is material to the examination of this application in accordance with Title 37, Code of Federal Regulations § 1.56(a). Such information is material when it is not cumulative to information already of record or being made of record in the application, and

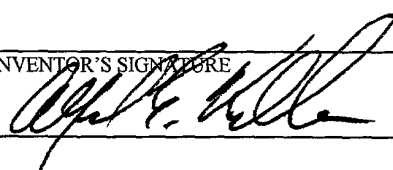
- (1) It establishes, by itself or in combination with other information, a prima facie case of unpatentability of a claim; or
- (2) It refutes, or is inconsistent with, a position the applicant has taken or may take in:
 - (I) opposing an argument of unpatentability relied on by the Office, or
 - (II) asserting an argument of patentability.

I hereby claim foreign priority benefits under Title 35, United States Code § 119 of any foreign application(s) for patent or inventor's certificates listed below and have also identified below any foreign application(s) having a filing date before that of the application(s) on which priority is claimed:

COUNTRY	APPLICATION NUMBER	DATE OF FILING	PRIORITY CLAIMED UNDER 35 USC 119

I hereby claim the benefit under Title 35 United States Code § 120 of any United States application(s) listed below and, insofar as any subject matter of any claim of this application is not disclosed in the prior United States Application, I acknowledge the duty to disclose material information as defined in Title 37, Code of Federal Regulations § 1.56(a) which occurred between the filing date of the prior application and the national PCT International filing date of this application: U.S. Provisional Application No. 60/146,635, filed July 30, 1999.

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

FULL NAME OF FIRST JOINT INVENTOR	INVENTOR'S SIGNATURE	DATE
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Same		

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE

Applicants: CONOCO INC. §
 Serial No.: Not Yet Assigned §
 Filed: Concurrently Herewith §
 For: Process for Producing Syngas in a Short §
 Contact Time Reactor Using Catalytic §
 Partial Oxidation of Hydrogen Sulfide §

POWER OF ATTORNEY BY ASSIGNEE

Atty. Dkt. No.: 1856-00301

Under the provisions of 37 C.F.R. § 3.71, the undersigned assignee of record of the entire interest in the above-identified patent/patent application by virtue of an assignment recorded (check as applicable):

☒ Concurrently herewith
☐ Date Recorded
☐ Reel _____ Frame _____

elects to conduct the prosecution of the application/maintenance of the patent to the exclusion of the inventor(s). The undersigned hereby declares that he has reviewed the above-referenced assignment and hereby declares that, to the best of his knowledge, title is in the Assignee and he is empowered to sign on behalf of the Assignee, and further declares that all statements made herein of his own knowledge are true and that all statements made on information and belief are believed to be true. The assignee hereby revokes any previous powers of attorney and appoints the following to prosecute this application/maintain this patent and transact all business in the Patent and Trademark Office connected therewith:

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ASSIGNEE

CONOCO INC.

Date: 21 JULY 2000

BY: Michael A. Gist
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 TITLE: CORPORATE SECRETARY